**1.2.2 EXERCISE 4 - ELECTRONEGATIVITY AND BOND POLARITY**

1. Explain the meaning of the term electronegativity.

2. Draw the correct dipoles on the following covalent bonds:

a) Cl – Cl

b) H – O

c) B – F

d) N – F

e) N – O

f) H – Cl

g) H – N

h) B – H

3. Draw the following molecules, add dipoles to all the bonds and predict whether or not they have an overall dipole:

a) CO2

b) SO2

c) BF3

d) NF3

e) Cl2

f) HCl