**5.3A HONORS HOMEWORK – INTRODUCTION TO OXIDATION AND REDUCTION**

1. Write half-equations to show the following changes, and indicate whether they represent oxidation or reduction:

|  |  |  |
| --- | --- | --- |
| Mg losing two electrons | Mg 🡪 Mg2+ + 2e | oxidation |
| Cl2 turning into 2Cl- | Cl2 + 2e- 🡪 2Cl- | reduction |
| Tl+ losing two electrons |  |  |
| 2H+ becoming H2 |  |  |
| V2+ losing one electron |  |  |
| Al3+ gaining three electrons |  |  |
| 2O2- becoming O2 |  |  |

1. Complete the following table to show the name and formula of some common ionic compounds:

|  |  |
| --- | --- |
| Name | Formula |
| magnesium oxide | MgO |
| iron (II) chloride | FeCl2 |
| lead (IV) oxide |  |
|  | PbO |
| lead (II) chloride |  |
|  | AlCl3 |
| iron (III) nitrate |  |
|  | FeSO4 |